

**Simple gravimetric chemical analysis –
weighing molecules the microscale way**

Activity 2: Student worksheet

Mass of hydrated copper(II) sulfate used = _____ g

Mass of anhydrous copper(II) sulfate after heating = _____ g

Mass of water removed by heating = _____ g

Number of moles of copper sulfate (CuSO_4) left at the end after water was removed
= mass of copper(II) sulfate/gram formula mass of copper (II) sulfate = _____

Number of moles of water removed by heating
= mass of water/gram formula mass of water = _____

Ratio of moles of water to copper sulfate = moles of H_2O /moles CuSO_4 = _____

The formula of hydrated copper(II) sulfate is $\text{CuSO}_4 \cdot \underline{\quad} \text{H}_2\text{O}$.