

Worksheet 1

Understanding precipitates – solubility rules

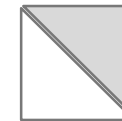
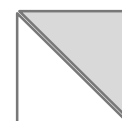
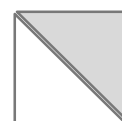
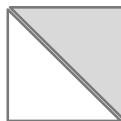
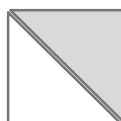
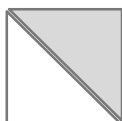
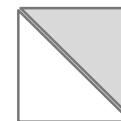
Add 2 drops of 0.1 M copper(II) sulfate over the diagonal in the square.

Add 2 drops of 0.4 M sodium hydroxide to the 'puddle'.

Stir with the pointed end of a wooden splint.

Disposal: wipe over the plastic surface with a paper towel.

In the squares below, you can try other reactions with available chemicals, e.g., silver nitrate with potassium iodide or sodium carbonate and magnesium sulfate (a 'green' reaction with chemicals available from the local shop).



Where do the components of the solutions come from?

In the small circle on the right, add some copper chloride crystals. Add water with a pipette to fill the whole circle and stir with a freshly cut pointed splint to dissolve the solid.



In the small circle on the left, add some anhydrous sodium carbonate. Add water with a pipette to fill the whole circle and stir with a freshly cut pointed splint to dissolve the solid.

With a pipette, move 2 drops of the liquid in the circle on the left into the square in the middle. With another pipette, move 2 drops of the liquid in the circle on the right into the square in the middle. Stir the contents of the square with a freshly cut pointed splint.